

AP Chemistry Pacing Guide

Units	Big Idea	Learning Objectives	Time
Unit 1: Atomic Structure and Properties - Moles & Molar Mass - Mass Spectroscopy of Elements - Elemental Composition of Pure Substances - Composition of Mixtures - Atomic Structure and Electron Configuration - Photoelectron Spectroscopy - Periodic Trends - Valance Electrons & Ionic Compounds	BIG IDEA 1: Scale, Proportion, and Quantity BIG IDEA 2: Structure and Properties	- I can calculate quantities of a substance or its relative number of particles using dimensional analysis and the mole concept. - I can explain the quantitative relationship between the mass spectrum of an element and the masses of the element's isotopes. - I can explain the quantitative relationship between the elemental composition by mass and the empirical formula of a pure substance. - I can explain the quantitative relationship between the elemental composition by mass and the composition of substances in a mixture. - I can represent the electron configuration of an element or ions of an element using the Aufbau principle. - I can explain the relationship between the photoelectron spectrum of an atom or ion and: a. The electron configuration of the species. b. The interactions between the electrons and the nucleus. - I can explain the relationship between trends in atomic properties of elements and electronic structure and periodicity. - I can explain the relationship between trends in the reactivity of elements and periodicity.	6 days
Unit 2: Molecular and Ionic Compound Structure and Properties - Types of Chemical Bonds - Intramolecular Force and Potential Energy - Structure of Ionic Solids - Structure of Metals & Alloys - Lewis Diagrams - Resonance & Formal Charge - VSEPR & Bond Hybridization	BIG IDEA 2: Structure and Properties	- I can explain the relationship between the type of bonding and the properties of the elements participating in the bond. - I can represent the relationship between potential energy and distance between atoms, based on factors that influence the interaction strength. - I can represent an ionic solid with a particulate model that is consistent with Coulomb's law and the properties of the constituent ions. - I can represent a metallic solid and/or alloy using a model to show essential characteristics of the structure and interactions present in the substance. - I can represent a molecule with a Lewis diagram. - I can represent a molecule with a Lewis diagram that accounts for resonance between equivalent structures or that uses formal charge to select between nonequivalent structures. - Based on the relationship between Lewis diagrams, VSEPR theory, bond orders, and bond polarities: - I can explain the structural properties of molecules.	6 days

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