

Reversible Reactions and Chemical Equilibrium Worksheet

Supporting questions: Answers

- (1) Can you draw the reversible arrow symbol? 
- (2) What do you observe when you heat up hydrated copper sulfate? *A colour change from blue to white and water vapour (steam) produced.*
- (3) Define the term dynamic equilibrium? *When the forward reaction occurs at the same rate as the backward reaction in a closed system.*
- (4) Why can dynamic equilibrium only occur in a closed system? *So there is no loss of products or reactants from the system.*
- (5) What are the 3 factors that alter the position of equilibrium? *Temperature, Concentration of reactants/products and pressure in gas.*
- (6) What happens to the position of equilibrium if you add a catalyst? *No change to the position of equilibrium it only speeds up the reaction so equilibrium is reached more quickly.*
- (7) For the chemical reaction below can you say what happens to the concentration of ammonia if you (a) add in more hydrogen (b) increase the pressure (c) increase the temperature (the forward reaction is exothermic). *(a) Equilibrium shifts to the right to produce more products. (b) Equilibrium shifts to the left to favour the endothermic reaction so less products. (c) Equilibrium shifts to the right as less molecules of gas so more products formed.*



- (8) Why might a high temperature still be used? *It is a compromise – less products will be formed but the rate is faster.*

Thanks for taking the time to complete these questions, you can check your answers on the next page.

Reversible Reactions And Equilibrium Worksheet

Thomas Bond, Chris Hughes



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