

• Angular Quantum Number (l)

- also known as azimuthal quantum number
- describes the shape of the orbital
- Orbitals have shapes that are best described as spherical ($l = 0$), polar ($l = 1$), or cloverleaf ($l = 2$). They can even take on more complex shapes as the value of the angular quantum number becomes larger.



• Magnetic Quantum Number (m_l)

- to describe the orientation in space of a particular orbital.
- It is called the magnetic quantum number because the effect of different orientations of orbitals was first observed in the presence of a magnetic field.

Rules Governing the Allowed Combinations of Quantum Numbers

1. The three quantum numbers (n , l , and m_l) that describe an orbital are integers: 0, 1, 2, 3, and so on.
2. The principal quantum number (n) cannot be zero. The allowed values of n are therefore 1, 2, 3, 4, and so on.
3. The angular quantum number (l) can be any integer between 0 and $n - 1$. If $n = 3$, for example, l can be either 0, 1, or 2.
4. The magnetic quantum number (m_l) can be any integer between $-l$ and $+l$. If $l = 2$, m_l can be either -2, -1, 0, +1, or +2.

Atomic Orbitals

- ψ is a mathematical function that describes the wave-like behavior of either one electron or a pair of electrons in an atom. This function can be used to calculate the probability of finding any electron of an atom in any specific region around the atom's nucleus. The term may also refer to the physical region or space where the electron can be calculated to be present, as defined by the particular mathematical form of the orbital.

- The second character identifies the subshell. By convention, the following lowercase letters are used to indicate different subshells.

s :	$l = 0$
p :	$l = 1$
d :	$l = 2$
f :	$l = 3$

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