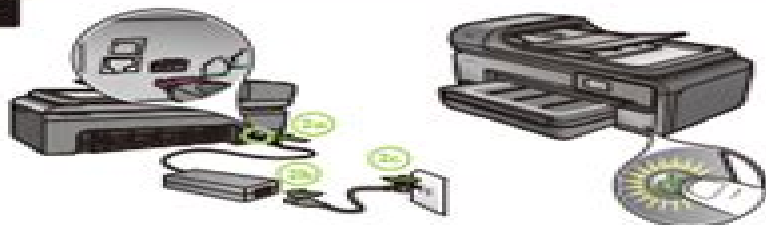




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Thomas Griffiths



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 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$
 $3\text{K}_2\text{SO}_4 + 4\text{H}_2\text{O} + 7\text{CH}_4 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O} + 8\text{H}_2\text{O} + 3\text{C}_3\text{H}_8 + 5\text{O}_2 + 2\text{CO}_2 + 4\text{H}_2\text{O}$

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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3$
 $\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 5\text{NF}_3 \rightarrow 2\text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 6\text{NaBr}$
 $3\text{SNO}_2 + \text{NF}_3 + 3\text{H}_2\text{O} \rightarrow 2\text{N}_2\text{O}_3$ using the following equation $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$
 $5\text{SNO}_2 + 3\text{NF}_3 \rightarrow \text{SnF}_2 + 2\text{N}_2\text{O}_3$
 using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$
 $5\text{SNO}_2 + 3\text{NF}_3 \rightarrow \text{SnF}_2 + 2\text{N}_2\text{O}_3$

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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$
 $5\text{SNO}_2 + 3\text{NF}_3 \rightarrow \text{SnF}_2 + 2\text{N}_2\text{O}_3$
 solve the following stoichiometry grams grams problems 6

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{HCl} \rightarrow \text{HCl}$ MgCl_2 H_2 O H_2O note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.41 \times 10^5 \text{ y}$ θ π $8.52 \times 10^9 \text{ s}$

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