

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

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Table of Contents Periodic Trends Answers Chemfiesta

1. Understanding the eBook Periodic Trends Answers Chemfiesta
 - The Rise of Digital Reading Periodic Trends Answers Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Answers Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Answers Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Answers Chemfiesta
 - Personalized Recommendations
 - Periodic Trends Answers Chemfiesta User Reviews and Ratings
 - Periodic Trends Answers Chemfiesta and Bestseller Lists
5. Accessing Periodic Trends Answers Chemfiesta Free and Paid eBooks

- Periodic Trends Answers Chemfiesta Public Domain eBooks
 - Periodic Trends Answers Chemfiesta eBook Subscription Services
 - Periodic Trends Answers Chemfiesta Budget-Friendly Options
6. Navigating Periodic Trends Answers Chemfiesta eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Answers Chemfiesta Compatibility with Devices
 - Periodic Trends Answers Chemfiesta Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Answers Chemfiesta
 - Highlighting and Note-Taking Periodic Trends Answers Chemfiesta
 - Interactive Elements Periodic Trends Answers Chemfiesta
 8. Staying Engaged with Periodic Trends Answers Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Answers Chemfiesta
 9. Balancing eBooks and Physical Books Periodic Trends Answers Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Answers Chemfiesta
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Trends Answers Chemfiesta
 - Setting Reading Goals Periodic Trends Answers Chemfiesta
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Trends Answers Chemfiesta
 - Fact-Checking eBook Content of Periodic Trends Answers Chemfiesta
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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