

Chapter 5: Periodic Law

- I. Worksheet on History of the Periodic Table and Periodic Law.
 - A. Dobereiner, Newlands, Mendeleev, Moseley.
 - B. **Periodic Law:** The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- II. The Periodic Table
 - A. **Periodic Table:** an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column or group.
 - B. Information in each block of the table:
 1. Symbol
 2. Atomic Number
 3. Atomic Mass
 4. Name
 5. Electron Configuration
 - C. The Periodic Table is organized into rows and columns of elements.
 1. **Periods:** Horizontal rows of elements. Numbered 1-7. Those numbers correspond to the principle energy level (n).
 2. **Groups:** Vertical columns of elements. Also called families. numbered in 3 ways.
 - a. **European System:** 1A-VIIIA; IB-VIII B
 - b. **U.S. System:** 1A, 2A, 3B-8B, 1B, 2B, 3A-8A. We will use this one.
 - c. **IUPAC System:** 1-18
 3. Family Names to know:

<u>Group</u>	<u>Family Name</u>
1A	Alkali Metals
2A	Alkaline Earth Metals
7A	Halogens
8A	Noble Gases

Other groups with A are called by the name of the first element in that group. For example, Group 4A is called the Carbon Group (Family).
 4. The elements with B in the group number are the transition metals.
 5. The Lanthanides and Actinides are on the bottom of the table. They are called the inner transition metals. They don't fit into the table very well. They make it too big.

Periodic Law Study Guide Answers

Jo A. Beran



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