Many neutralization reactions are exothermic. In part 3 we will determine the heat of neutralization of hydrochloric acid with sodium hydroxide according to the reaction

$$HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H2O(l)$$

This reaction takes place in the same calorimeter as above and we can summarize the heat flow the same way as we do in part 2: $-q_{exn} = q_{suter} + q_{calorimeter}$

Finally, we will measure the heat of fusion of ice in part 4. To accomplish that we add ice, initially at 0 °C, to the calorimeter with warm or room temperature water. The warmth of the water will be distributed to the

Part 3: Enthalpy of Neutralization Reaction

Combine about 25 mL of 3M NaOH with about 25 mL of 3 M HCl. Record volumes and temperatures.

Constant Constant Constant	Trial 1	Trial 2	Trial 3
Volume of 3M HCl	24.0 mL	25.3 mL	25.8 mL
Volume of 3M NaOH	25.1 mL	27.1 mL	26.4 mL
T _i of solution	21.5 °C	22.6 °C	23.1 °C
Tf of solution	34.0 °C	33.2 °C	34.8 °C
ΔH _{rxn} in J/g			1-0

Average ΔH _{rxn} in J/g:	
Average ΔH _{ren} in kJ/mol:	
Average arrest in Karmon	

The standard enthalpy change for this reaction is -57.0 kJ/mol. Determine percent error in the experiment.

Percent	error:	
The second second		

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- 1. Understanding the eBook Neutralisation Exothermic Reaction Results Volume
 - The Rise of Digital Reading Neutralisation Exothermic Reaction Results Volume
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Neutralisation Exothermic Reaction Results Volume
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Neutralisation Exothermic Reaction Results Volume
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Neutralisation Exothermic Reaction Results Volume
 - Personalized Recommendations
 - Neutralisation Exothermic Reaction Results Volume User Reviews and Ratings
 - Neutralisation Exothermic Reaction Results Volume and Bestseller Lists
- 5. Accessing Neutralisation Exothermic Reaction Results Volume Free and Paid eBooks
 - Neutralisation Exothermic Reaction Results Volume Public Domain eBooks
 - Neutralisation Exothermic Reaction Results Volume eBook Subscription Services
 - Neutralisation Exothermic Reaction Results Volume Budget-Friendly Options

- 6. Navigating Neutralisation Exothermic Reaction Results Volume eBook Formats
 - o ePub, PDF, MOBI, and More
 - Neutralisation Exothermic Reaction Results Volume Compatibility with Devices
 - Neutralisation Exothermic Reaction Results Volume Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Neutralisation Exothermic Reaction Results Volume
 - Highlighting and Note-Taking Neutralisation Exothermic Reaction Results Volume
 - Interactive Elements Neutralisation Exothermic Reaction Results Volume
- 8. Staying Engaged with Neutralisation Exothermic Reaction Results Volume
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Neutralisation Exothermic Reaction Results Volume
- 9. Balancing eBooks and Physical Books Neutralisation Exothermic Reaction Results Volume
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Neutralisation Exothermic Reaction Results Volume
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Neutralisation Exothermic Reaction Results Volume
 - Setting Reading Goals Neutralisation Exothermic Reaction Results Volume
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Neutralisation Exothermic Reaction Results Volume
 - Fact-Checking eBook Content of Neutralisation Exothermic Reaction Results Volume
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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