

Chemistry 100 Final Study Guide Key

• Chapter 1: Introduction

a. Physical vs. Chemical properties

1) Physical properties: shiny, metallic, malleable, ductile, conductive. Chemical property: reactive with water.

b. Units of measurement

1) Meters, liters, grams (with any prefixes) are metric system. All others are U.S.

2) 5.68×10^{-7} 46,300,000

c. Density

1) 2.47g/mL

• Chapter 3: Atomic Structure

a. Historical Perspective

1) List the major scientific accomplishment of the following scientists:

(i) Experiments in electrochemistry described ions and the electrical nature of atoms

(ii) Cathode-ray tube experiments led to the proposal of the "Plum pudding" model of atoms

(iii) Oil-drop experiments calculated the mass of electrons

(iv) Gold foil experiments proved atoms had a dense, positively-charged nucleus and empty space

(v) Proposed an atomic model with a nucleus and electrons circling like planets

(vi) Proposed that electrons exist in a "cloud" or probability region around the nucleus

b. Isotopes

1) Isotopes of an element have the same number of protons, but different numbers of neutrons. Regular hydrogen has 1 proton and no neutrons. Isotopes could have 1 proton and 1 neutron (deuterium) or 1 proton and 2 neutrons (tritium).

c. Structure of the atom

1) Protons and neutrons are in the nucleus. Electrons are in regions of space surrounding the nucleus.

2) 3 energy levels, 4 sublevels, 6 orbitals, 11 electrons

3) $1s^2 2s^2 2p^6 3s^2 3p^2$

d. Electron configurations and trends on the Periodic Table

1) See the periodic Table.

• Chapter 5: Chemical bonds

a. Electron-dot structures

1) Draw the electron-dot structure of an oxygen atom.



b. Octet Rule and Ions

1) Mg^{2+} , Cl^- , N^{3-}

c. Ionic compounds

1) Aluminum bromide, calcium sulfide, sodium nitride

d. Covalent compounds

1) Sulfur dioxide, phosphorus trichloride, carbon tetrahydride

e. Polyatomic molecules

1) Nitrate (NO_3^-), chlorate (ClO_3^-), carbonate (CO_3^{2-}), sulfate (SO_4^{2-}), chromate (CrO_4^{2-}), phosphate (PO_4^{3-})

f. Electron-dot formulas for molecules

1) Answer will be given in class.

g. VSEPR Model and Molecular shape



VSEPR shape = pyramidal

Principles Of Chemistry Final Study Guide

Robert Truscott



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Student's Guide to Fundamentals of Chemistry Jo A. Beran, 2013-10-22 *Student's Guide to Fundamentals of Chemistry* Fourth Edition provides an introduction to the basic chemical principles. This book deals with various approaches to chemical principles and problem solving in chemistry. Organized into 25 chapters, this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry. This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for the elements. Other chapters consider how to calculate the molecular weight of a compound from its formula. This book discusses as well the characteristics of a photon in terms of its particle-like properties and defines the wavelength, frequency, and speed of light. The final chapter deals with the fundamental components of air and the classification of materials formed in natural waters. This book is a valuable resource for chemistry students, lecturers, and instructors.

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