

Name:

Period:

Stoichiometry Unit Study Guide

Vocabulary section

Directions: Use the following words from the word bank to fill in the blanks that best fits the sentence/definitions. Each word is only used **ONCE**.

Word bank			
Stoichiometry	Grams	Molar Mass	Mole ratio
Mole number	Scientific notation	Law of conservation of mass	Balanced chemical equation

1. The _____ states that in a chemical reaction, mass is **not** created or destroyed. It also means that if an element is present among the reactants it must almost be present in the product.
2. _____ is the mass of one mole of any atom/molecule and has a unit of **g/mol**.
3. _____ is the unit that measures mass only.
4. _____ is the relationship between the amounts/quantities of substances taking part in a chemical reaction. We use the "train track method" to solve these types of problems.
5. The _____ is **1 mole = 6.02×10^{23} atoms/molecules**.
6. A _____ is like the **recipe** of a chemical reaction and must follow the law of conservation.
7. _____ is a way of writing very large or very small numbers to use them more effectively.
8. _____ is a **conversion factor** that relates the amounts in moles of any two substances in a chemical reaction.

Directions: Answer the following questions using dimensional analysis (AKA the train track method). Show all your work and include units!

1. Find the number of moles in 46.32 grams of H_2O .

Section Study Guide Stoichiometry Answers

James E. Brady, Fred Senese



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Chemistry, Student Study Guide** James E. Brady, Fred

Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane CH_4 from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. *Study Guide to Accompany Calculus for the Management, Life, and Social Sciences* Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management, Life, and Social Sciences. *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even-numbered end-of-chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

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