

Section 1

SECTION 1

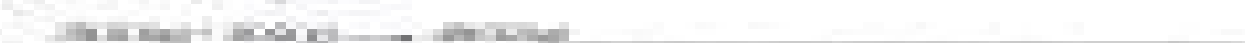
Answer each question on the following questions in the space provided.

1. ☐ a. The following are chemical equations representing the reaction of hydrogen gas with oxygen gas.
(a) $2H_2 + O_2 \rightarrow 2H_2O$
(b) $2H_2 + O_2 \rightarrow 2H_2O$
(c) $2H_2 + O_2 \rightarrow 2H_2O$
(d) $2H_2 + O_2 \rightarrow 2H_2O$
(e) $2H_2 + O_2 \rightarrow 2H_2O$
(f) $2H_2 + O_2 \rightarrow 2H_2O$
(g) $2H_2 + O_2 \rightarrow 2H_2O$
(h) $2H_2 + O_2 \rightarrow 2H_2O$
(i) $2H_2 + O_2 \rightarrow 2H_2O$
(j) $2H_2 + O_2 \rightarrow 2H_2O$
(k) $2H_2 + O_2 \rightarrow 2H_2O$
(l) $2H_2 + O_2 \rightarrow 2H_2O$
(m) $2H_2 + O_2 \rightarrow 2H_2O$
(n) $2H_2 + O_2 \rightarrow 2H_2O$
(o) $2H_2 + O_2 \rightarrow 2H_2O$
(p) $2H_2 + O_2 \rightarrow 2H_2O$
(q) $2H_2 + O_2 \rightarrow 2H_2O$
(r) $2H_2 + O_2 \rightarrow 2H_2O$
(s) $2H_2 + O_2 \rightarrow 2H_2O$
(t) $2H_2 + O_2 \rightarrow 2H_2O$
(u) $2H_2 + O_2 \rightarrow 2H_2O$
(v) $2H_2 + O_2 \rightarrow 2H_2O$
(w) $2H_2 + O_2 \rightarrow 2H_2O$
(x) $2H_2 + O_2 \rightarrow 2H_2O$
(y) $2H_2 + O_2 \rightarrow 2H_2O$
(z) $2H_2 + O_2 \rightarrow 2H_2O$
2. ☐ b. The following are chemical equations representing the reaction of hydrogen gas with oxygen gas.
(a) $2H_2 + O_2 \rightarrow 2H_2O$
(b) $2H_2 + O_2 \rightarrow 2H_2O$
(c) $2H_2 + O_2 \rightarrow 2H_2O$
(d) $2H_2 + O_2 \rightarrow 2H_2O$
(e) $2H_2 + O_2 \rightarrow 2H_2O$
(f) $2H_2 + O_2 \rightarrow 2H_2O$
(g) $2H_2 + O_2 \rightarrow 2H_2O$
(h) $2H_2 + O_2 \rightarrow 2H_2O$
(i) $2H_2 + O_2 \rightarrow 2H_2O$
(j) $2H_2 + O_2 \rightarrow 2H_2O$
(k) $2H_2 + O_2 \rightarrow 2H_2O$
(l) $2H_2 + O_2 \rightarrow 2H_2O$
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(n) $2H_2 + O_2 \rightarrow 2H_2O$
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(x) $2H_2 + O_2 \rightarrow 2H_2O$
(y) $2H_2 + O_2 \rightarrow 2H_2O$
(z) $2H_2 + O_2 \rightarrow 2H_2O$
3. ☐ c. The following are chemical equations representing the reaction of hydrogen gas with oxygen gas.
(a) $2H_2 + O_2 \rightarrow 2H_2O$
(b) $2H_2 + O_2 \rightarrow 2H_2O$
(c) $2H_2 + O_2 \rightarrow 2H_2O$
(d) $2H_2 + O_2 \rightarrow 2H_2O$
(e) $2H_2 + O_2 \rightarrow 2H_2O$
(f) $2H_2 + O_2 \rightarrow 2H_2O$
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(v) $2H_2 + O_2 \rightarrow 2H_2O$
(w) $2H_2 + O_2 \rightarrow 2H_2O$
(x) $2H_2 + O_2 \rightarrow 2H_2O$
(y) $2H_2 + O_2 \rightarrow 2H_2O$
(z) $2H_2 + O_2 \rightarrow 2H_2O$

4. Write the balanced chemical equation for the following reaction:



a. Balance the equation:



b. Write the balanced chemical equation for the reaction of H_2 and O_2 .

c. Write the balanced chemical equation for the reaction of H_2 and O_2 to form H_2O .

d. Write the balanced chemical equation for the reaction of H_2 and O_2 to form H_2O .

e. Write the balanced chemical equation for the reaction of H_2 and O_2 to form H_2O .

Section 1 Review Stoichiometry

David Baud



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