

Student Information

SECTION 1

Answer each question by circling the correct response(s) in the space provided.

1. The following is a chemical equation representing the reaction of hydrogen gas with oxygen gas to form water.

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$
 - (a) The reaction is exothermic.
 - (b) The reaction is endothermic.
 - (c) The reaction is spontaneous.
 - (d) The reaction is non-spontaneous.
2. The following is a chemical equation representing the reaction of hydrogen gas with oxygen gas to form water.

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$
 - (a) The reaction is exothermic.
 - (b) The reaction is endothermic.
 - (c) The reaction is spontaneous.
 - (d) The reaction is non-spontaneous.
3. The following is a chemical equation representing the reaction of hydrogen gas with oxygen gas to form water.

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$
 - (a) The reaction is exothermic.
 - (b) The reaction is endothermic.
 - (c) The reaction is spontaneous.
 - (d) The reaction is non-spontaneous.
4. The following is a chemical equation representing the reaction of hydrogen gas with oxygen gas to form water.

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$
 - (a) The reaction is exothermic.
 - (b) The reaction is endothermic.
 - (c) The reaction is spontaneous.
 - (d) The reaction is non-spontaneous.

5. Write the balanced chemical equation for the following reaction:



a. Balance the equation:



b. Write the balanced chemical equation for the reaction of hydrogen gas with oxygen gas to form water.

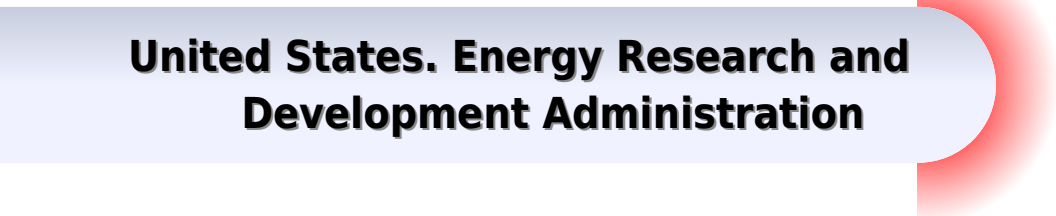
c. Write the balanced chemical equation for the reaction of hydrogen gas with oxygen gas to form water.

d. Write the balanced chemical equation for the reaction of hydrogen gas with oxygen gas to form water.

e. Write the balanced chemical equation for the reaction of hydrogen gas with oxygen gas to form water.

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**United States. Energy Research and
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